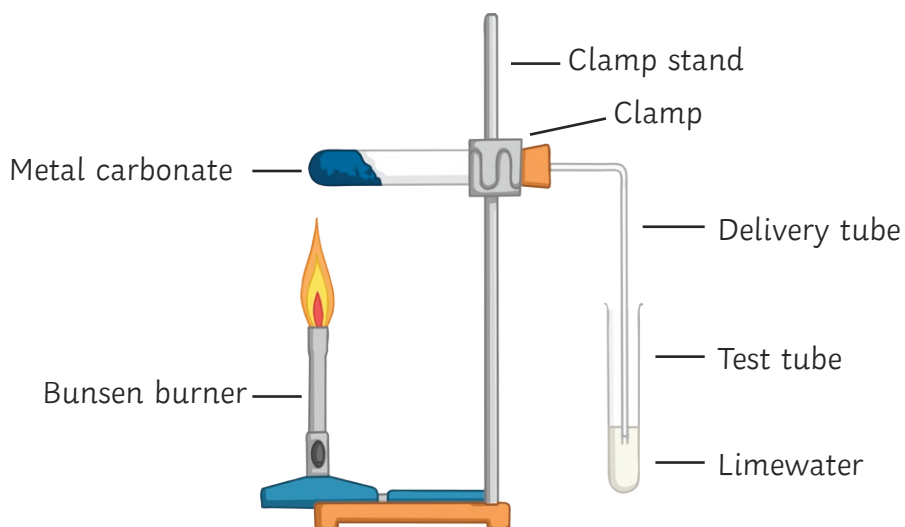


Chemical Reactions

Exam Style Questions 2

1. Thermal decomposition of copper carbonate



- State what is meant by thermal decomposition.
- What effect would adding a catalyst have on the experiment?
- Complete the equation.

copper carbonate \rightarrow _____ + carbon dioxide

- Complete the table below to show the mass of carbon dioxide produced when 120g of copper carbonate is broken down completely.

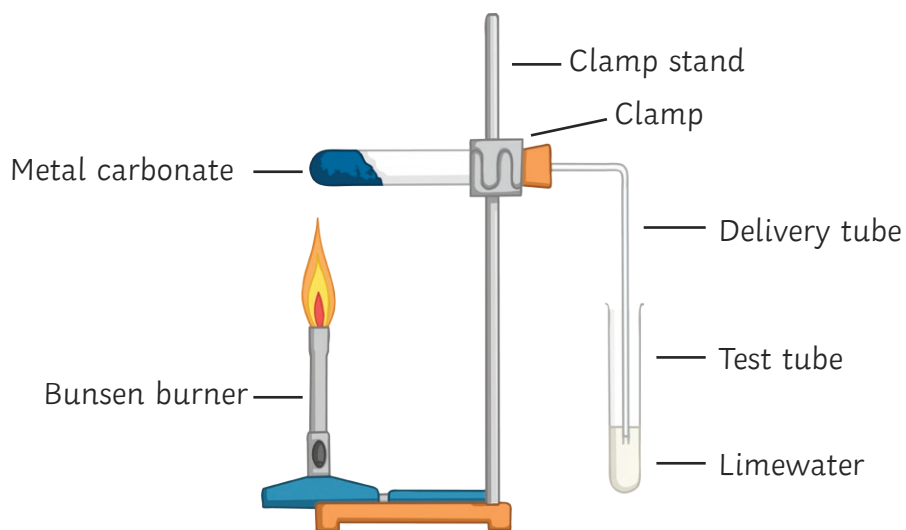
Compound	Reactant or product	Mass (g)
Copper carbonate	Reactant	120
Unnamed in equation	Product	75
Carbon dioxide	Product	

2. State 2 changes that may occur in a chemical reaction
3. State 1 difference between an exothermic and endothermic reaction

Chemical Reactions

Exam Style Questions 2 Answers

1. Thermal decomposition of copper carbonate



- State what is meant by thermal decomposition.
when heat causes the breakdown of a substance
- What effect would adding a catalyst have on the experiment?
Catalysts speed up the rate of reaction
- Complete the equation.
copper carbonate \rightarrow **copper oxide** + carbon dioxide
- Complete the table below to show the mass of carbon dioxide produced when 120g of copper carbonate is broken down completely.

Compound	Reactant or product	Mass (g)
Copper carbonate	Reactant	120
Unnamed in equation	Product	75
Carbon dioxide	Product	45

- State 2 changes that may occur in a chemical reaction
Temperature rise, colour change, bubbles of gas, flames
- State 1 difference between an exothermic and endothermic reaction
Exothermic reactions show an increase in temperature, endothermic reactions show a decrease in temperature